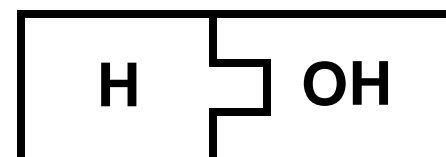
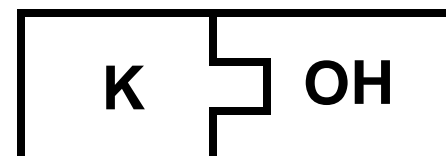
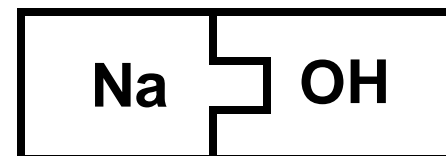
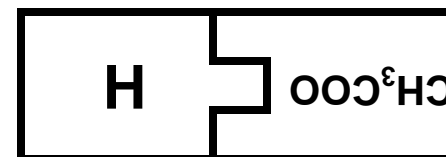
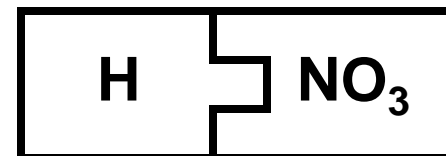
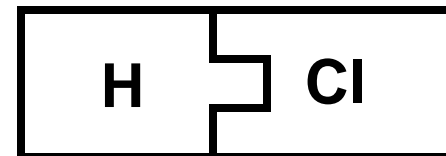
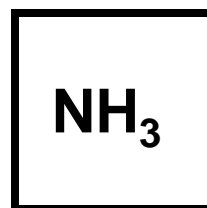
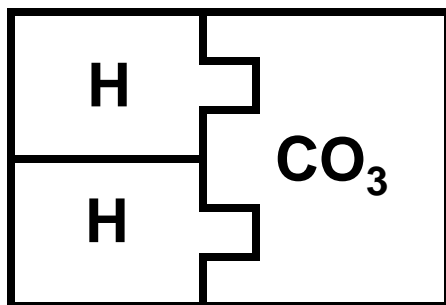
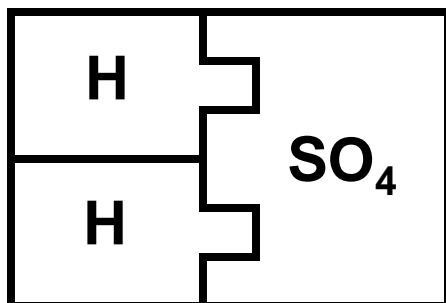
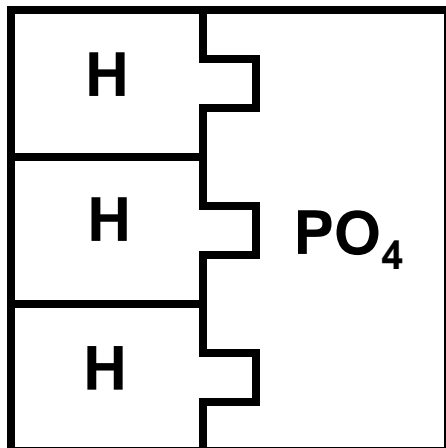


# Chemical Formula Models - Acids, Bases and Polyatomic Ions



# The Chemical Formula of Acids, Bases and Polyatomic Ions

## 1. Acids

Phosphoric acid	$\text{H}_3\text{PO}_4$
Sulfuric acid	$\text{H}_2\text{SO}_4$
Carbonic acid	$\text{H}_2\text{CO}_3$
Hydrochloric acid	$\text{HCl}$
Nitric acid	$\text{HNO}_3$
Acetic acid	$\text{CH}_3\text{COOH}$

Acids produce hydrogen ions in solution,  $\text{H}^+$

## 2. Alkalis (soluble bases)

Sodium hydroxide	$\text{NaOH}$
Potassium hydroxide	$\text{KOH}$
Ammonia	$\text{NH}_3$

Bases produce hydroxide ions in solution,  $\text{OH}^-$

## 3. Ions (charged particles)

→	Phosphate ion	$\text{PO}_4^{3-}$
→	Sulfate ion	$\text{SO}_4^{2-}$
→	Carbonate ion	$\text{CO}_3^{2-}$
→	Chloride ion	$\text{Cl}^-$
→	Nitrate ion	$\text{NO}_3^-$
→	Acetate ion	$\text{CH}_3\text{COO}^-$
	Hydrogen ion	$\text{H}^+$
→	Hydroxide ion	$\text{OH}^-$
→	Hydroxide ion	$\text{OH}^-$
→	Ammonium ion	$\text{NH}_4^+$

## 4. Chemical Equations

