

Balancing Chemical Equations Worksheet 1

- $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
- $\text{Na} + \text{Cl}_2 \rightarrow \text{NaCl}$
- $\text{N}_2\text{O}_4 \rightarrow \text{NO}_2$
- $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$
- $\text{H}_2\text{O}_2 \rightarrow \text{H}_2\text{O} + \text{O}_2$
- $\text{Ca} + \text{N}_2 \rightarrow \text{Ca}_3\text{N}_2$
- $\text{Li} + \text{F}_2 \rightarrow \text{LiF}$
- $\text{Mg} + \text{N}_2 \rightarrow \text{Mg}_3\text{N}_2$
- $\text{NH}_3 \rightarrow \text{N}_2 + \text{H}_2$
- $\text{HCl} \rightarrow \text{H}_2 + \text{Cl}_2$
- $\text{NI}_3 \rightarrow \text{N}_2 + \text{I}_2$
- $\text{HI} \rightarrow \text{H}_2 + \text{I}_2$
- $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
- $\text{K} + \text{S} \rightarrow \text{K}_2\text{S}$
- $\text{Li} + \text{Cl}_2 \rightarrow \text{LiCl}$
- $\text{HgO} \rightarrow \text{Hg} + \text{O}_2$
- $\text{Al} + \text{I}_2 \rightarrow \text{AlI}_3$
- $\text{Al} + \text{S} \rightarrow \text{Al}_2\text{S}_3$
- $\text{Na} + \text{H}_2\text{O} \rightarrow \text{NaOH} + \text{H}_2$
- $\text{NO} + \text{O}_2 \rightarrow \text{NO}_2$
- $\text{Na}_2\text{O} + \text{H}_2\text{O} \rightarrow \text{NaOH}$
- $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- $\text{PbO} + \text{C} \rightarrow \text{Pb} + \text{CO}_2$
- $\text{C} + \text{O}_2 \rightarrow \text{CO}$
- $\text{CO} + \text{O}_2 \rightarrow \text{CO}_2$

Balancing Chemical Equations Worksheet 2

26. $\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$
27. $\text{Ag}_2\text{O} \rightarrow \text{Ag} + \text{O}_2$
28. $\text{K} + \text{O}_2 \rightarrow \text{K}_2\text{O}$
29. $\text{Cl}_2 + \text{F}_2 \rightarrow \text{ClF}_3$
30. $\text{SiO}_2 + \text{C} \rightarrow \text{Si} + \text{CO}$
31. $\text{NaHCO}_3 \rightarrow \text{Na}_2\text{CO}_3 + \text{CO}_2 + \text{H}_2\text{O}$
32. $\text{ZnS} + \text{O}_2 \rightarrow \text{ZnO} + \text{SO}_2$
33. $\text{SO}_2 + \text{H}_2\text{S} \rightarrow \text{S} + \text{H}_2\text{O}$
34. $\text{Ba} + \text{O}_2 \rightarrow \text{BaO}$
35. $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$
36. $\text{SO}_2 + \text{O}_2 \rightarrow \text{SO}_3$
37. $\text{KIO}_3 \rightarrow \text{KI} + \text{O}_2$
38. $\text{NOCl} \rightarrow \text{NO} + \text{Cl}_2$
39. $\text{COF}_2 \rightarrow \text{CO}_2 + \text{CF}_4$
40. $\text{Fe}_2\text{O}_3 + \text{Al} \rightarrow \text{Al}_2\text{O}_3 + \text{Fe}$
41. $\text{P}_4 + \text{O}_2 \rightarrow \text{P}_2\text{O}_5$
42. $\text{Na}_3\text{N} \rightarrow \text{Na} + \text{N}_2$
43. $\text{P}_4 + \text{Cl}_2 \rightarrow \text{PCl}_3$
44. $\text{CrO}_3 \rightarrow \text{Cr}_2\text{O}_3 + \text{O}_2$
45. $\text{P}_2\text{O}_5 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{PO}_4$
46. $\text{NaH} + \text{BF}_3 \rightarrow \text{B}_2\text{H}_6 + \text{NaF}$
47. $\text{P}_4 + \text{O}_2 \rightarrow \text{P}_4\text{O}_{10}$
48. $\text{BN} + 3\text{IF} \rightarrow \text{NI}_3 + \text{BF}_3$
49. $\text{Fe}_2\text{O}_3 + \text{C} \rightarrow \text{Fe} + \text{CO}_2$
50. $\text{Fe}_3\text{O}_4 + \text{CO} \rightarrow \text{Fe} + \text{CO}_2$

Balancing Chemical Equations Worksheet 1 - ANSWERS

- $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$
- $\text{N}_2\text{O}_4 \rightarrow 2\text{NO}_2$
- $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$
- $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$
- $3\text{Ca} + \text{N}_2 \rightarrow \text{Ca}_3\text{N}_2$
- $2\text{Li} + \text{F}_2 \rightarrow 2\text{LiF}$
- $3\text{Mg} + \text{N}_2 \rightarrow \text{Mg}_3\text{N}_2$
- $2\text{NH}_3 \rightarrow 3\text{N}_2 + 2\text{H}_2$
- $2\text{HCl} \rightarrow \text{H}_2 + \text{Cl}_2$
- $2\text{NI}_3 \rightarrow \text{N}_2 + 3\text{I}_2$
- $2\text{HI} \rightarrow \text{H}_2 + \text{I}_2$
- $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
- $\text{K} + 2\text{S} \rightarrow \text{K}_2\text{S}$
- $2\text{Li} + \text{Cl}_2 \rightarrow 2\text{LiCl}$
- $2\text{HgO} \rightarrow 2\text{Hg} + \text{O}_2$
- $2\text{Al} + 3\text{I}_2 \rightarrow 2\text{AlI}_3$
- $2\text{Al} + 3\text{S} \rightarrow \text{Al}_2\text{S}_3$
- $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$
- $2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2$
- $\text{Na}_2\text{O} + \text{H}_2\text{O} \rightarrow 2\text{NaOH}$
- $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
- $2\text{PbO} + \text{C} \rightarrow 2\text{Pb} + \text{CO}_2$
- $2\text{C} + \text{O}_2 \rightarrow 2\text{CO}$
- $2\text{CO} + \text{O}_2 \rightarrow 2\text{CO}_2$

Balancing Chemical Equations Worksheet 2 - ANSWERS

26. $2\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$
27. $2\text{Ag}_2\text{O} \rightarrow 4\text{Ag} + \text{O}_2$
28. $4\text{K} + 2\text{O}_2 \rightarrow 2\text{K}_2\text{O}$
29. $\text{Cl}_2 + 3\text{F}_2 \rightarrow 2\text{ClF}_3$
30. $\text{SiO}_2 + 2\text{C} \rightarrow \text{Si} + 2\text{CO}$
31. $2\text{NaHCO}_3 \rightarrow \text{Na}_2\text{CO}_3 + \text{CO}_2 + \text{H}_2\text{O}$
32. $2\text{ZnS} + 3\text{O}_2 \rightarrow 2\text{ZnO} + 2\text{SO}_2$
33. $\text{SO}_2 + 2\text{H}_2\text{S} \rightarrow 3\text{S} + 2\text{H}_2\text{O}$
34. $2\text{Ba} + \text{O}_2 \rightarrow 2\text{BaO}$
35. $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
36. $2\text{SO}_2 + \text{O}_2 \rightarrow 2\text{SO}_3$
37. $2\text{KIO}_3 \rightarrow 2\text{KI} + 3\text{O}_2$
38. $2\text{NOCl} \rightarrow 2\text{NO} + \text{Cl}_2$
39. $2\text{COF}_2 \rightarrow \text{CO}_2 + \text{CF}_4$
40. $\text{Fe}_2\text{O}_3 + 2\text{Al} \rightarrow \text{Al}_2\text{O}_3 + 2\text{Fe}$
41. $\text{P}_4 + 5\text{O}_2 \rightarrow 2\text{P}_2\text{O}_5$
42. $2\text{Na}_3\text{N} \rightarrow 6\text{Na} + \text{N}_2$
43. $\text{P}_4 + 6\text{Cl}_2 \rightarrow 4\text{PCl}_3$
44. $4\text{CrO}_3 \rightarrow 2\text{Cr}_2\text{O}_3 + 3\text{O}_2$
45. $\text{P}_2\text{O}_5 + 3\text{H}_2\text{O} \rightarrow 2\text{H}_3\text{PO}_4$
46. $6\text{NaH} + 2\text{BF}_3 \rightarrow \text{B}_2\text{H}_6 + 6\text{NaF}$
47. $\text{P}_4 + 5\text{O}_2 \rightarrow \text{P}_4\text{O}_{10}$
48. $\text{BN} + 3\text{IF} \rightarrow \text{NI}_3 + \text{BF}_3$
49. $2\text{Fe}_2\text{O}_3 + 3\text{C} \rightarrow 4\text{Fe} + 3\text{CO}_2$
50. $\text{Fe}_3\text{O}_4 + 4\text{CO} \rightarrow 3\text{Fe} + 4\text{CO}_2$